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POGIL: Periodic Trends in Atomic Properties Information Properties such as the size of an atom (atomic radius), the energy required to remove an electron from an atom (ionization energy), and the energy associated with the addition of an electron to a gaseous atom (electron affinity) can be understood in terms of the electron configuration of the atom and a competition between electron-nucleus attraction and electron-electron repulsion.

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[POGIL: Periodic Trends in Atomic Properties](#)

It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions.

[More on Configurations and Periodic Trends Why?](#)

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Across the periodic table, sizes of atoms show the following trends, with many irregularities: Size increases down a group. The outermost electrons are in successively more extensive orbitals as n increases.

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Periodic Trends POGIL Model 1 - Main Group Elements The periodic table is often considered to be the "best friend" of chemists and chemistry students alike. ... Support your answer, using examples from two periods.

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Periodic Trends 3 6. Locate the numbers in Model 1 that represent the ionization energy. The ionization energy is the amount of energy needed to remove an electron from an atom. a. Using your knowledge of Coulombic attraction, explain why ionization—removing an electron from an atom—takes energy.

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SWBAT recognize and explain the reason for the periodic trends of atomic radius, ionization energy, and electronegativity. Big Idea Coulombic attraction and atomic structure are responsible for properties of elements, which determines their arrangement on the periodic table.

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