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The chapters in this section deal with the rates and mechanisms of chemical change. These topics stand in contrast to the subjects of equilibrium and thermodynamics that control the direction of chemical change.. The Chem1 Virtual Textbook is a resource for General Chemistry aimed mainly at the first-year university level.

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Chemical Kinetics Factors That Affect Reaction Rates • Physical State of the Reactants In order to react, molecules must come in contact with each other. If the reaction is happening between a solid and a liquid it will react only on the surface. The more homogeneous the mixture of reactants, the faster the molecules can react.

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Presents a balanced presentation of the macroscopic view of empirical kinetics and the microscopic molecular viewpoint of chemical dynamics. Stressing interconnections between phenomenological chemical kinetics and molecular reaction dynamics, the book discusses reactions in gas phase, liquids, and at surfaces; molecular potential surfaces; gas-gas and gas-surface theories applied to reactive ...

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The kinetics fundamentals we covered in the earlier sections of this lesson group relate to processes that take place in the gas phase. But chemists and biochemists are generally much more concerned with solutions.

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Chemical kinetics, also known as reaction kinetics, is the branch of physical chemistry that is concerned with understanding the rates of chemical reactions. It is to be contrasted with thermodynamics, which deals with the direction in which a process occurs but in itself tells nothing about its rate.

~~Chemical Kinetics and Reaction Dynamics, Paul L. Houston~~

Chemical Kinetics and Reaction Dynamics brings together the major facts and theories relating to the rates with which chemical reactions occur from both the macroscopic and microscopic point of view. This book helps the reader achieve a thorough understanding of the principles of chemical kinetics

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Overview. Researchers in chemical physics apply physical methods and theory to study molecular and collective properties of chemical systems. The focus is on understanding complex phenomena from gas phase molecular dynamics, to nanoscale, mesoscale, and biological phenomena, through model systems and fundamental physical principles.

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both the specialist in the field and upper undergraduate and graduate-level chemistry students. It bridges the gap between the two with a path that leads the reader from phenomenological approach, to rates of chemical reactions, and then to the state-of-the-art calculation of the rate constants of ...

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